

Synthesis, structure and properties of V(V) monooxido complex with ONO tridentate Schiff base

Anna Jurowska, Janusz Szklarzewicz*, Maciej Hodorowicz, Ryszard Gryboś

Jagiellonian University, Faculty of Chemistry, Gronostajowa 2, 30-387 Kraków, Poland

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Abstract

The oxidovanadium(V) Schiff base complex of formula $[\text{VO}(\text{L})(\text{EtO})(\text{EtOH})]$ (where H_2L = Schiff base ligand derived from 5-methoxysalicylaldehyde and phenylacetic hydrazide) was synthesized and described. Complex crystallizes in triclinic P_1 space group. Octahedral geometry of the vanadium(V) centre is filled with oxido, ONO L^{2-} ligand and two solvent molecules both in ethoxo and as neutral ethanol form. The complex is neutral, with 5- and 6-membered ring formed by ONO ligand coordinated in octahedral plane with oxido and EtOH ligands in vertical positions. Two isomers are present in the unit cell, with different position of 5-membered ring *versus* vertical plane. The elemental analysis, magnetic susceptibility, thermogravimetry and spectroscopy (IR, UV-Vis) measurements were measured and are discussed. The cyclic voltammetry measurements show irreversible processes for vanadium(IV/V) redox system. Thermal stability both in a solid state (TG and SDTA measurements) as well as in solutions (at pH 7.0 and 2.0, studied by UV-Vis spectroscopy) is discussed.

Keywords: vanadium, complex, Schiff base, structure, 5-methoxysalicylaldehyde, phenylacetic hydrazide

Introduction

Vanadium(V) forms stereochemically flexible complexes, where the thermodynamically acceptable coordination geometries range from tetrahedral and octahedral to trigonal pyramidal and pentagonal bipyramidal [1]. The vanadium-containing compounds are particularly interesting due to their use as insulin-mimetic agents and there are three general classes of such compounds: 1) inorganic vanadium anionic or cationic salts ($[\text{VO}_4]^{3-}$ or VO^{2+}), 2) chelated vanadium(IV) complexes and 3) complexes of vanadium(V) like mono- and dioxovanadates $\{[\text{VO}(\text{O}_2)(\text{H}_2\text{O})_2(\text{L}_2)]^{n-}$ ($n = 0, 1$) and $[\text{VO}(\text{O}_2)_2(\text{L}_2)]^{n-}$ ($n = 1, 2, 3$), where L_2 denote bidentate ligand} [2]. The discovery of oxovanadates(V) and oxovanadium(IV) with the insulin-like *in vivo* and *in vitro* activity, as for example VOSO_4 , can have a promising use in the treatment of non-insulin-dependent type-2 diabetes mellitus [3,4]. Certain vanadium complexes like as bis(maltolato)oxovanadium(IV), bis(ethylmaltolato)oxovanadium(IV) or bis(isopropylmaltolato)oxovanadium(IV) are found to be promising candidates as oral complements for insulin in the treatment of diabetes [5-8]. In comparison to VOSO_4 the mentioned vanadium complexes have less toxicity, higher potency and better absorption in human digestive system [5]. The search for new vanadium complex compounds with decreased toxicity and increased potency still continues. The vanadium

complexes with Schiff base ligands have also become promising [9-11].

In this paper we described the synthesis and physicochemical characterization of new vanadium(V) complex with Schiff base ligand derived from 5-methoxysalicylaldehyde and phenylacetic hydrazide for which the determination of stability in solutions at pH 7.0 and 2.0, studied by UV-Vis spectroscopy, allowed to determine the possibility of using this compound as potential insulin-mimetic agent.

Materials and Methods

$[\text{VO}(\text{acac})_2]$, 5-methoxysalicylaldehyde and phenylacetic hydrazide, were of analytical grade (Aldrich) and were used as supplied. Ethanol (98%) of pharmaceutical grade was from Polmos, all other solvents were of analytical grade and were used as supplied. BaSO_4 was of spectroscopic grade (Japan). Bu_4NPF_6 was synthesized from Bu_4NBr and KPF_6 by a standard method [12] and recrystallized from acetone. Microanalysis of carbon, hydrogen and nitrogen were performed using Elementar Vario MICRO Cube elemental analyzer. IR spectrum was recorded on a Bruker EQUINOX 55 FT-IR spectrophotometer in KBr pellets. The electronic absorption spectra were recorded on Shimadzu UV-3600 UV-Vis-NIR spectrophotometer equipped with a CPS-240 temperature controller. Complex stability measurements in $\text{DMSO}/\text{H}_2\text{O}$ (20 $\mu\text{l}/3\text{ml}$) mixture were performed at $c = 5.0 \times 10^{-5}$ M, pH = 7.00 and at pH = 2.00 (0.1M KCl, HCl added to adjust the pH). The spectra were measured every 340 s

*Corresponding author: szklarze@chemia.uj.edu.pl,
tel: 48-12-686-2504, fax: 48-12-686-2750

at 37 °C on Shimadzu UV-3600 UV-Vis-NIR spectrophotometer in 1 cm UV cuvettes. The magnetic susceptibility measurement was performed on a SHERWOOD SCIENTIFIC magnetic susceptibility balance. The cyclic voltammetry measurements were carried out in DMSO with $[\text{Bu}_4\text{N}]\text{PF}_6$ (0.1 M) as the supporting electrolyte, using Pt working and counter and Ag/AgCl as reference electrodes on an AUTOLAB/PGSTAT 128 N Potentiostat/Galvanostat. $E_{1/2}$ values were calculated from the average anodic and cathodic peak potentials, $E_{1/2} = 0.5(E_a + E_c)$. The redox potentials were calibrated *versus* ferrocene (0.440 V *versus* SHE), which was used as an internal potential standard for measurements in organic solvents to avoid the influence of a liquid junction potential; the final values are reported *versus* the standard hydrogen electrode (SHE). DTG measurements were performed on a TGA/SDTA 851e Mettler Toledo Microthermogravimeter under argon (in the 25-1000 °C range) and in air atmosphere (in the 25-750 °C range) with scan speed 10 °C/min.

Synthesis

In the one step reaction, the Schiff base ligand (L) was synthesized *in situ* from 5-methoxysalicylaldehyde and phenylacetic hydrazide in 1 : 1 molar ratio. Then, the vanadyl acetylacetonate was added under anaerobic conditions (Ar).

$[\text{VO}(\text{L})(\text{EtO})(\text{EtOH})]$, **1**

The 5-methoxysalicylaldehyde (0.187 ml, 1.5 mmol) and the phenylacetic hydrazide (0.225 g, 1.5 mmol) in 50 ml of ethanol were refluxed for 11 minutes under Ar. Then the $[\text{VO}(\text{acac})_2]$ (0.399 g, 1.5 mmol) was added. The reflux was continued for 58 minutes under Ar. The solution was evaporated in part but precipitation was observed next day in refrigerator. The obtained very dark green (almost black) compound was filtered off, washed with ethanol and dried in air. Yield: 0.410 g, 62 %. MW = 440.36. Anal. Calcd. for $\text{C}_{20}\text{H}_{25}\text{N}_2\text{O}_6\text{V}$: C, 54.55; H, 5.72; N, 6.36 %. Found: C, 54.36; H, 5.64; N, 6.34 %. The complex is diamagnetic.

Crystallographic data collection and structure refinement

The crystal of **1** suitable for X-ray analysis was selected from the materials prepared as described in the Synthetic procedure. The crystal data, details of data collection and structure refinement parameters are summarized in Table 1. Intensity data for **1** was collected on a Super Nova diffractometer using monochromatic MoK α radiation, $\lambda = 0.71073$ Å. The structure was solved by direct methods [13] and the least-squares refinements were performed with the SHELXL-2014 program [14]. The positions of all atoms were determined by direct methods. All non-hydrogen atoms were refined anisotropically using weighted full-matrix least-squares on F^2 . Hydrogens were included in the structure at idealized positions. The figures were made using DIAMOND software [15]. Crystallographic data for the structural analysis

have been deposited with the Cambridge Crystallographic Data Centre, CCDC No. 1871642.

Results and Discussion

The reaction of *in situ* formed Schiff base ligand based on 5-methoxysalicylaldehyde and phenylacetic hydrazide with $[\text{VO}(\text{acac})_2]$ in ethanolic solution, under anaerobic conditions, results in formation of green-black transparent solution. Cooling the solution enabled crystals to be obtained. The magnetic susceptibility measurement indicates on V(V) oxidation state for complex **1**. The compound is stable in air and soluble in organic solvents, but almost insoluble in water.

1. Crystal structure

The complex **1** crystallizes in the triclinic space group P-1 with the asymmetric cell unit containing two molecules of $[\text{VO}(\text{L})(\text{EtO})(\text{EtOH})]$ complex. The selected bond lengths and angles for **1** are listed in Table 2, whereas the molecular structure is presented in Fig. 1. The Schiff base ligand coordinates in a tridentate mode to afford a distorted octahedral geometry. In complex **1** the O7, N11 and O21 from the L_{ONO^2} Schiff base ligand and oxygen atom (O22) of EtO⁻ ligand are occupying equatorial, whereas oxido ligand (O28) and oxygen atom (O25) of EtOH ligand axial positions. The V=O bond is the shortest [1.5837(15) Å], while V-O25 (from EtOH molecule) is the longest one, as expected for the *trans* effect [2.3824(17) Å]. The vanadium atom is shifted of the square plane, occupied by ONO ligand and EtO⁻ molecule, by 0.304 Å. The phenolic ring is strongly bent, the angle between plane with salicylaldehyde and that with phenolic plane is 73.69 °. The similar angle was observed in other com-

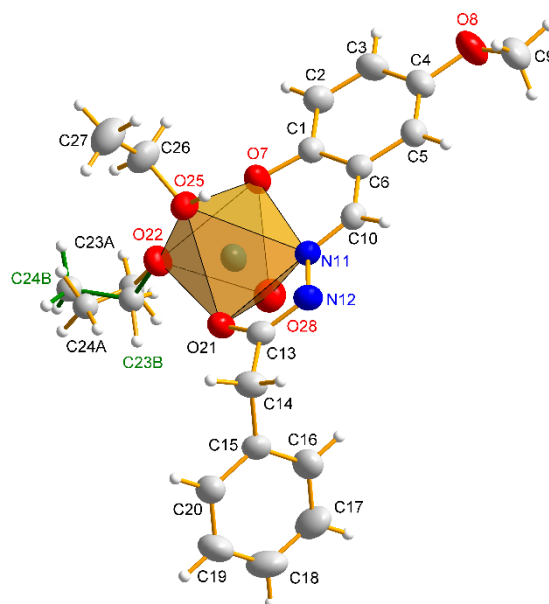


Figure 1. The molecular structure of complexes **1** together with the atom labelling scheme. All non-hydrogen atoms are represented at 30% probability thermal ellipsoids

Table 1. Crystal data and structure refinement for **1**

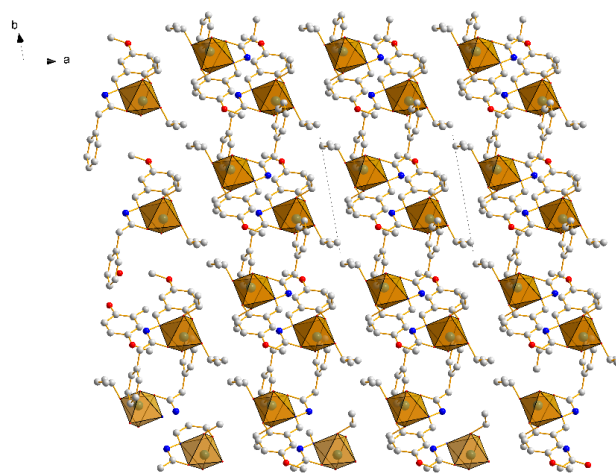
	1
Empirical formula	C ₂₀ H ₂₅ N ₂ O ₆ V
Formula weight	440.36
Crystal size (mm)	0.630 x 0.190 x 0.100
Crystal system	Triclinic
Space group	P -1
a (Å)	10.467(5)
b (Å)	10.683(5)
c (Å)	10.991(5)
α (°)	118.752(5)
β (°)	94.900(5)
γ (°)	96.272(5)
h	-13 ≤ h ≤ 13
k	-13 ≤ k ≤ 13
l	-13 ≤ l ≤ 13
V (Å ³)	1057.5(9)
Z	2
T (K)	293(2)
Wavelength [Å]	0.71073
D _x (mg/m ³)	1.383
Absorption coefficient (mm ⁻¹)	0.507
F(000)	460
Theta range for data collection (°)	3.722 to 26.374
Reflections collected	13522
Independent reflections	4316 [R(int) = 0.0224]
Completeness to theta	99.7 % (25.242°)
Absorption correction	Semi-empirical from equivalents
Refinement method	Full-matrix least-squares on F ²
Data / restraints / parameters	4316 / 24 / 289
Goodness-of-fit on F ²	1.082
Final R indices [I > 2σ(I)]	R1 = 0.0343, wR2 = 0.0942
R indices (all data)	R1 = 0.0456, wR2 = 0.0989
Largest diff. peak and hole (e. Å ⁻³)	0.239 and -0.234

Table 2. Selected bond lengths [Å] and bond angles [°] in **1**

bond lengths		bond angles	
V(1)-O(28)	1.5837(15)	O(28)-V(1)-O(22)	102.78(7)
V(1)-O(22)	1.7570(14)	O(28)-V(1)-O(7)	100.42(8)
V(1)-O(7)	1.8488(14)	O(22)-V(1)-O(7)	102.93(6)
V(1)-O(21)	1.9662(14)	O(28)-V(1)-O(21)	96.91(7)
V(1)-N(11)	2.1278(16)	O(22)-V(1)-O(21)	93.13(7)
V(1)-O(25)	2.3824(17)	O(7)-V(1)-O(21)	153.02(6)
C(10)-N(11)	1.287(2)	O(28)-V(1)-N(11)	95.06(6)
N(11)-N(12)	1.399(2)	O(22)-V(1)-N(11)	159.33(7)
N(12)-C(13)	1.305(2)	O(7)-V(1)-N(11)	83.74(6)
C(13)-O(21)	1.295(2)	O(21)-V(1)-N(11)	74.23(6)
O(22)-C(23B)	1.423(4)	O(28)-V(1)-O(25)	174.76(6)
O(22)-C(23A)	1.72(5)	O(22)-V(1)-O(25)	81.05(6)
O(25)-C(26)	1.436(2)	O(7)-V(1)-O(25)	82.03(6)
O(25)-H(25)	0.744(19)	O(21)-V(1)-O(25)	79.19(6)
C(26)-C(27)	1.492(3)	N(11)-V(1)-O(25)	80.54(5)
C(23A)-C(24A)	1.240(4)	C(1)-O(7)-V(1)	134.01(11)
C(23B)-C(24B)	1.399(10)	C(10)-N(11)-V(1)	127.57(12)
		N(12)-N(11)-V(1)	115.51(10)
		C(13)-O(21)-V(1)	118.41(11)
		C(23B)-O(22)-V(1)	129.9(3)
		C(23A)-O(22)-V(1)	128.5(12)
		C(26)-O(25)-V(1)	119.38(12)
		V(1)-O(25)-H(25)	116.6(16)

plexes with similar diphenyl ligands as for example for square planar copper complex – 70.81 °, for octahedral copper – 78.23 ° [16], for silver complex – 85.38 ° [17], for hexameric manganese complex – 68.80 ° [18] and for trimeric Ni(II) complex - 80.84 ° [19]. There are two coordination geometries of ONO ligand in 1 : 1 ratio present in the crystal, the planes containing two phenolic rings of two isomers are parallel to each other with the distance of 1.117 Å.

The molecular packing is presented in Fig. 2. Two adjacent complexes are connected by two hydrogen bonds between O25-N12, and O25'-N12' as shown in Fig. 3 and in Table 3. The bond distance is 2.825 Å, with the angle of 172 ° indicating on relatively strong interactions. Second hydrogen bond between O26 and O22 is very weak, as the angle is only 111.9 °. There are none π - π interactions or any other strong intermolecular interactions which could keep the structure and which could form a net of intermolecular interactions. Between layers, shown in Fig. 2, there are only short contacts. All this indicates that the structure is simply a packing of separated hydrogen bonded dimers.

**Figure 2.** The crystal packing in **1** through [001] direction**Table 3.** The hydrogen bonds for **1** [Å and °]

D-H...A	d(D-H)	d(H...A)	d(D...A)	<(DHA)
C(26)-H(26A)...O(22)	0.97	2.62	3.113(3)	111.9
O(25)-H(25)...N(12)#1	0.744(19)	2.09(2)	2.825(2)	172(2)

Symmetry transformations used to generate equivalent atoms: #1 -x+1,-y+1,-z+1

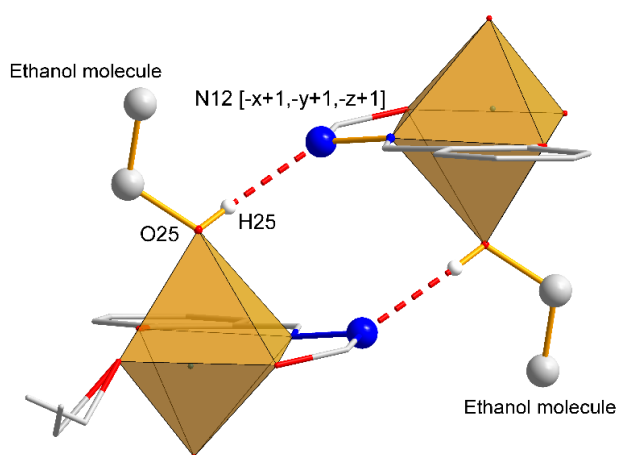


Figure 3. The hydrogen bonds in **1**

2. Spectral data

The IR spectrum of complex **1** is presented in Fig. 4. In the 900 – 1000 cm^{-1} range, characteristic for $\text{V}=\text{O}$ vibration, two intensive bands at 918 and 978 cm^{-1} are observed, which indicates on the V(V) oxidation state and it stays in line with magnetic susceptibility measurement. The band located at 1602 cm^{-1} can be assigned to $\text{C}=\text{N}$ group of Schiff base ligand. This band was also observed in the other oxido vanadium Schiff base complexes [20–22]. The lack of sharp band for ν_{OH} group at ca. 3500 cm^{-1} indicates on involving of this group in coordination to the metal center and it stays in agreement with the crystal structure results. The sharp band at 2973 cm^{-1} can be assigned to EtOH ligand in complex **1**.

The UV-Vis spectra were measured in water and in seven different organic solvents (Fig. 5). In the UV part of the electronic spectra the band attributed to the coordinated Schiff base ligand are observed as the separate peaks. In the visible part of the spectra, the band located at ca. 410 nm, can be assigned to a LMCT (ligand-to-metal charge-transfer) transition between metal center and Schiff base ligand. This band is clearly visible for spectra measured in acetonitrile and DMF. The molar absorption coefficient in DMSO is $5.06 \cdot 10^3 \text{ mol}^{-1} \cdot \text{cm}^{-1} \cdot \text{dm}^3$ confirming its CT character. The lack of bands above 500 nm confirms V(V) oxidation state and d^0 configuration.

The obtained complex is sparingly soluble in water and well soluble in DMF and DMSO. As, for biological activity, the solutions of the complex have to be delivered to *in vitro* and *in vivo* study, essential was to use aqueous solutions. Thus the concentrated solutions of **1** in DMSO were used and mixed with water prior to such a study. Therefore the stability of DMSO solutions of complex was studied as first and it was shown that in period of days DMSO solutions are stable in time. We did not use DMF, as DMSO is serviceable in for example medicine due to the fact that it supports the transfer of substances through biological membranes. The stability of the complex **1** was measured also in DMSO-water mixtures of composition used in biological tests. We used UV-Vis spectroscopy for stability investigations and DMSO- H_2O mixture (20 μl + 3ml, respectively) at neutral pH and at pH = 2.00. This last pH was used, as oral administration of vanadium complexes results in its passing through stomach (where pH changes in range 1.5–2.5) and it is important that complex can be still stable in such conditions,

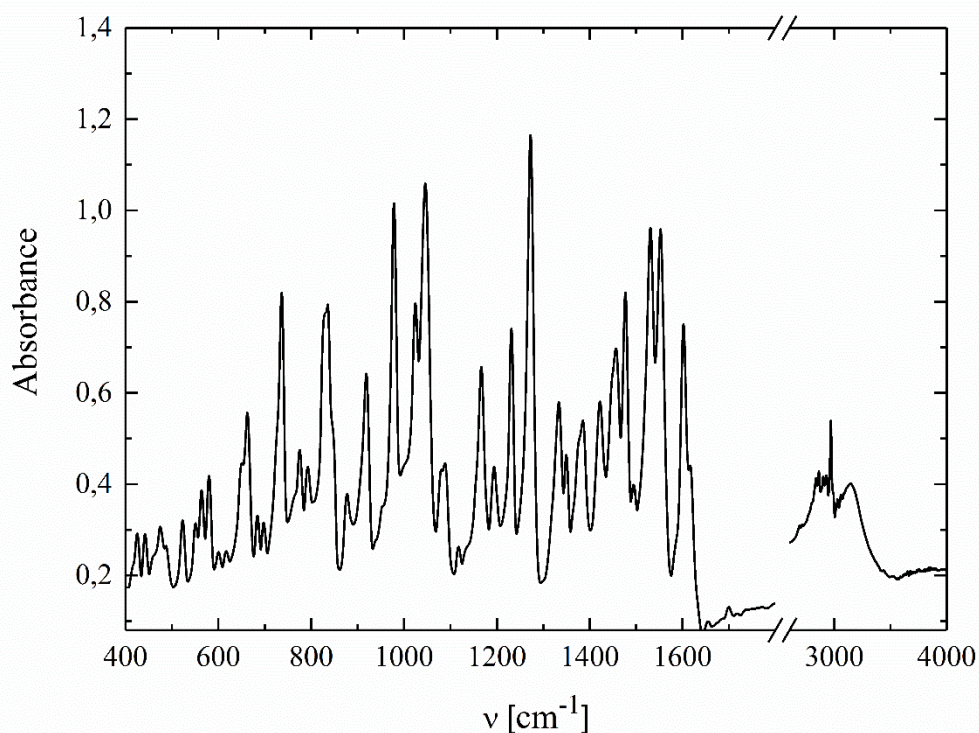


Figure 4. The IR spectrum of **1** in KBr

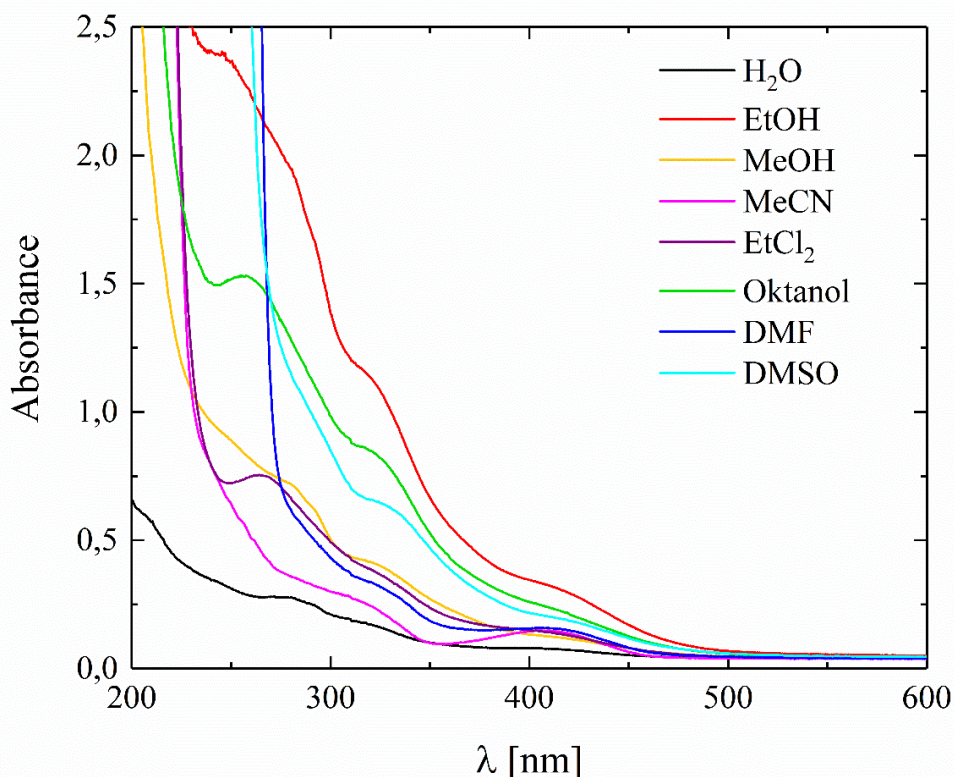


Figure 5. UV-Vis qualitative spectra of **1** in different solvents, $d = 1$ cm

at least for some time. The UV-Vis spectra were measured over time for each pH. The results of the measurements for $\text{pH} = 7.00$ are presented in Fig. 6 and for $\text{pH} = 2.00$ in Fig 7. The complex **1** is stable in neutral pH, but at $\text{pH} = 2.00$ it gradually and very slowly decomposes. The decrease of CT band at ca. 400 nm, indicates on ONO ligand release. The more complicated is the

interpretation of the spectral change at $\text{pH} = 2.0$. Even for the first measured spectrum, no band at ca. 400 nm is observed. This indicates, that Schiff base was completely removed from the complex within ca. one minute prior to spectroscopic measurements. This situation was also observed for the other vanadium Schiff base complex [20].

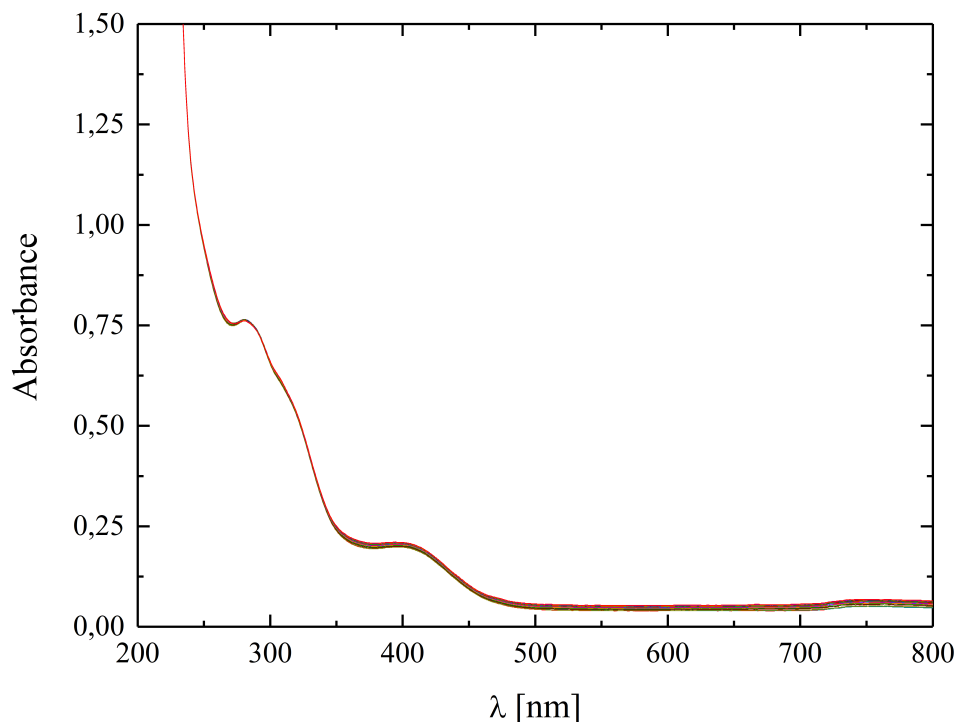


Figure 6. UV-Vis spectra of **1** in DMSO- H_2O (20 μl + 3 ml) mixture at $\text{pH} = 7.00$. $T = 37^\circ\text{C}$, spectra measured in 340 s intervals

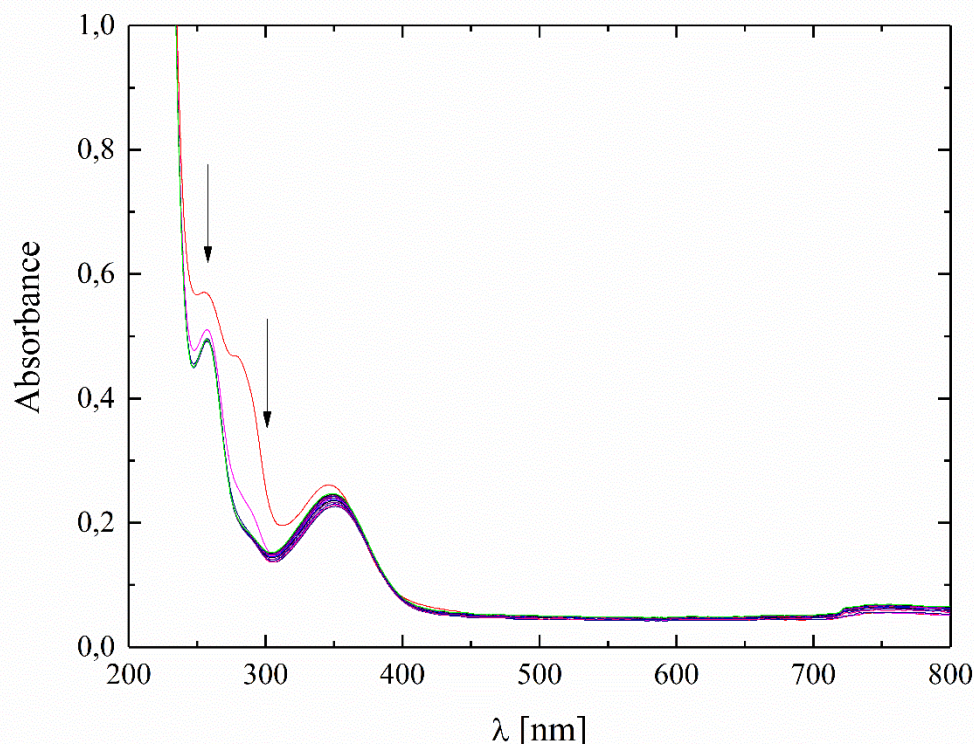


Figure 7. UV-Vis spectra of **1** in DMSO-H₂O (20 µl + 3 ml) mixture at pH = 2.00. T = 37 °C, spectra measured in 340 s intervals. The arrows indicate the direction of the changes

3. Cyclic voltammetry measurements

The cyclic voltammetry for complex **1** was performed in DMSO (Fig. 8). On the voltammetric curve, one anode peak can be distinguished with extremely high potential (0.430 V), which can be attributed to irreversible oxidation to V(V). Most likely, the

oxidized form is so unstable that it is almost completely consumed as a result of subsequent reactions. The peak position is independent on scan speed, as shown in inset of the Fig. 8. What is more, in the return scan, appear the secondary three cathode peaks (at 0.162, -0.291 and -0.708 V), corresponding to the processes of reduction of unstable anode reaction products.

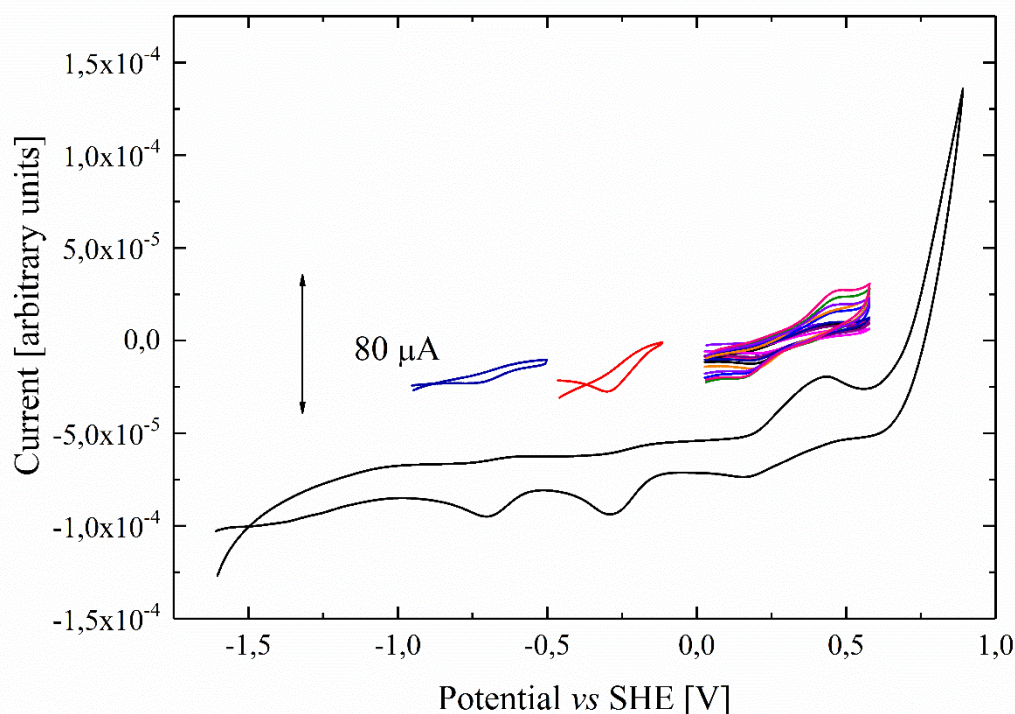


Figure 8. The cyclic voltammetry measurements of **1** in DMSO. Pt working and counting, Ag/AgCl as a reference electrode, 0.1M Bu₄NPF₆ as electrolyte. Scan speed 100 mV·s⁻¹. Insets, scan speed 100 mV·s⁻¹, for measurements at different scan speeds it vary from 20 to 250 mV s⁻¹

4. Thermogravimetry measurements

TG and SDTA measurements were performed in air or under argon with a heating rate 10°/min and are presented in Fig. 9 and Fig. 10 respectively. The numerical data are collected in Table 4.

L release. This last process is the most egzothermic one with SDTA T_{\max} at 524 °C. The ligand L is decomposed in two not resolved steps, indicating that it decomposes probably into initial aldehyde and hydrazide parts which are released in two distinct processes.

Table 4. TG and SDTA data for complex **1**. Heating rate 10°/min

Conditions	M [mg]	SDTA T_{\max} [°C]	process	Δm_{\exp} [%]	Δm_{calc} [%]	Release of
In air	8.94001	73.8	endo			
	8.0935	158	egzo	9.47	10.4	EtOH
	7.41445	524	egzo	17.1	20.8	EtOH
	6.33795			29.1		
	1.85378			79.3	85.0	L + 2EtOH
Argon	5.71728	69	endo			
	5.15267	167	endo	9.88	10.4	EtOH
	4.70634	315	endo	17.7		
	2.88585	362	endo	49.5	45	EtOH + aldehyde
	1.49052	451	endo	73.9	74.6	L + EtOH
		672	endo			
		848	endo			

The first process, observed for both experiments below 100 °C, is connected with the release of EtOH molecule in endothermic reaction. The lack of intermolecular hydrogen bonds, shown by X-ray measurements, and very weak intermolecular interaction allows the EtOH molecules to be released at relatively low temperature. There are two separated steps for EtOH release, at lower temperature coordinated neutral ethanol is released, while at higher one, second molecule of coordinated EtO^- is released. The second step in air is connected not only with the second EtOH molecule release, but this process overlaps with the ligand

Under Ar, all decomposition processes are endothermic. After one EtOH molecule release (as in air) the decomposition processes overlaps yielding not so well defined plateaus on TG curve and stepwise mass loose up to 1000 °C is observed. But even at this temperature no plateau on TG curve is observed. The final mass loose corresponds well with L and EtOH molecule release, but probably the scheme of complex decomposition is similar as in air, but due to the lack of oxygen, the vanadium carbides are formed at the end of the reaction decreasing the observed mass loose so much that part of the carbons from ligand L remains.

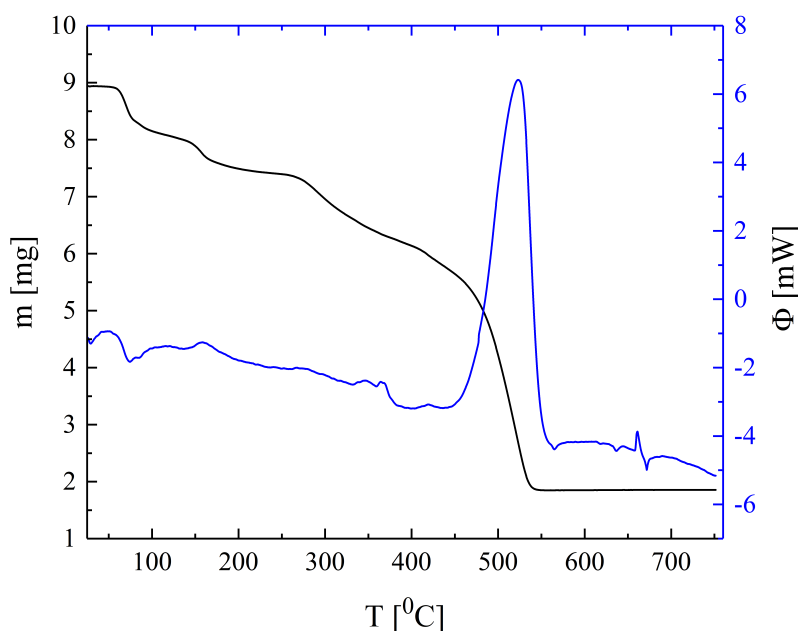


Figure 9. TG and SDTA curves of **1** in air. Scan rate 10°/min, air. The blue line represents the first derivative of TG curve

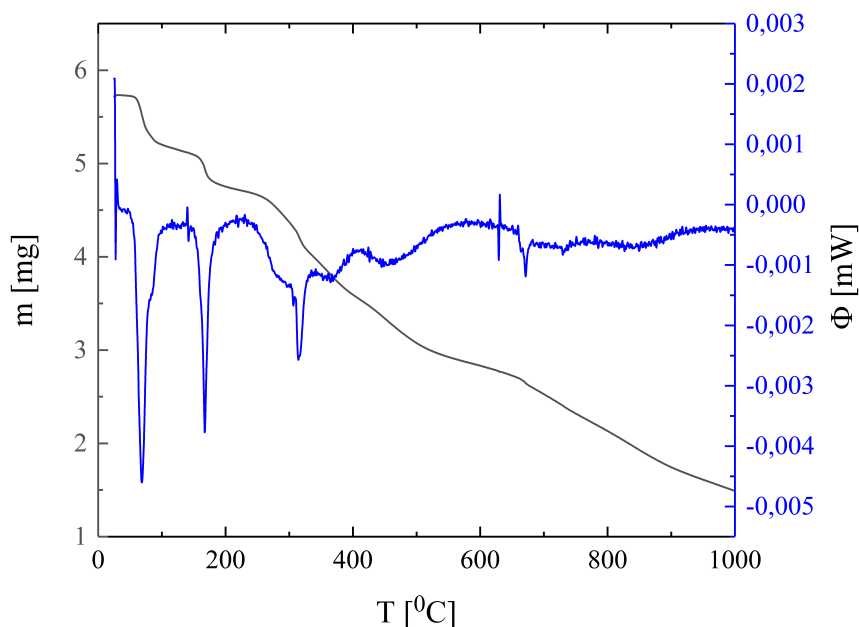


Figure 10. TG and SDTA curves of **1** under Ar. Scan rate 10°/min, air. The blue line represent the first derivative of TG curve

Conclusions

The reaction of Schiff base components (5-methoxysalicylaldehyde and phenylacetic hydrazide) with $[\text{VO}(\text{acac})_2]$ results in isolation of $[\text{VO}(\text{L})(\text{EtO})(\text{EtOH})]$, where L = ONO Schiff base ligand. The complex was structurally and spectroscopically characterized. The vanadium complexes with tridentate Schiff base ligand compared to bidentate complexes as for example bis(maltolato)oxovanadium(IV) (BMOV) seem to be more stable, thus they can probably be transferred to biological systems in unchanged form. Complex **1**, in water-DMSO mixture, gradually and very slowly decomposes, but it seems that most of the complex can survive in digestion systems, and this creates the possibility of using it as an insulin mimetic compound.

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